

General Chemistry I - Chemistry 1211 Fall 2008

Instructor: Dr. Carole Brown
208 Rogers Hall
cebrown@ngcsu.edu
864-1506

Office Hours: Monday, Wednesday, Thursday, and Friday at 11:00-12:00 Tuesday 10:00-12:00 and welcome by appointment

Lecture: 10:10am-11:05am MWF in 204 Rogers Hall
3 credit hours

Lab: Dependent on section registered for. *You must attend your scheduled day/time for both Lab and Lecture*

Co-requisites: Chemistry 1211L Laboratory

Co-requisites Policy: Although 1211 and 1211L are separate courses with separate grades, you are required to be enrolled in both. These classes are meant to compliment each other and reinforce the ideas presented in both courses. Laboratory information will be included on exams and homework.

Textbooks and Required Materials:

Lecture: McMurry & Fay, *Chemistry*, Prentice Hall, 5th Ed.

Bundled with Mastering Chemistry and a partial Solutions Manual

Laboratory: Modular Laboratory Program in Chemistry, ed. H.A. Neidig

***Make sure you get the correct packet

Carbonless-copy producing notebook

General: Non-programmable, single-line or double-line display, calculator with scientific functions (log, ln, 10^x, e^x, and EE or EXP)

Course Description and Objectives: The first course in a two-semester sequence covering the fundamental principles and applications of chemistry for science majors. Topics to be covered include composition of matter, dimensional analysis stoichiometry, periodic relations, and nomenclature. By the end of the semester, students will have learned and will be able to exhibit skill in the fundamental principles of chemistry these topics.

Lecture Attendance: REGULAR LECTURE ATTENDANCE IS ESSENTIAL TO YOUR SUCCESS IN THIS CLASS AND IS, THEREFORE, MANDATORY. In accordance with NGCSU policy, a student can be dropped for excessive absenteeism. If absent, you are responsible for determining what assignments, material, etc. you have missed.

If NGCSU officially cancels class due to inclement weather, any exam will be postponed until the next class meeting.

1211 Grading:

Exam 1	16%
Exam 2	16%
Exam 3	16%
Exam 4	16%
Homework	16%
Final Exam	20%

Note on Absences and Exam Policy: ANY UNEXCUSED ABSENCE FROM AN EXAM RESULTS IN A ZERO. FOR AN ABSENCE TO BE EXCUSED IT MUST BE DISCUSSED WITH ME IN PERSON AND APPROVED PRIOR TO THE ABSENCE. No additional time will be allotted to those late to any exam

Homework: Homework assignments can be found at MasteringChemistry.com. You must enroll as soon as possible using the course id MCBROWN31451.

You will be given an assignment at the beginning of each chapter. You will have a short period of time to complete these assignments, approximately 1 week. Late submissions will receive no credit.

Cheating and Plagiarism:

Any attempt at cheating or plagiarism will not be tolerated and will result in being failed in the course, reported to the V.P. for Academic affairs, the Dean of Students and the Academic Integrity Council, which may ultimately result in suspension or expulsion. See the Academic Integrity section below.

Class Evaluations: Class evaluations at NGCSU are conducted on-line through Banner. Evaluation of the class is considered a component of the course and students will not be permitted to access their course grade until the evaluation has been completed. The evaluations will be accessible beginning one week prior to Final Exam week.

Disabilities: NGCSU is committed to equal access to its programs, services, and activities for people with disabilities. If you will need accommodations in this class, reasonable prior notice needs to be given to the instructor and the office of Student Disability Resources. Please contact Ms. Elizabeth McIntosh, Coordinator, at 867-2782 in 122 Barnes Hall.

Early Intervention Program (EIP): As your teacher, I am committed to your academic success, not only in this class, but at this university. If I feel you would benefit from some of the special services available to students, I will make the appropriate referral. I will expect to comply with the referral and take advantage of the services offered.

Calculators: Calculators must be non-programmable and generally have a single-line display. The calculator must have scientific functions – such as log, ln, and EXP or EE. ALL OTHER TYPES OF CALCULATORS ARE PROHIBITED AND WILL BE CONFISCATED FOR THE DURATION OF THE TEST IF USED. There will be no passing or sharing of the calculators

Cell Phones: Use of your cell phone during lab, lecture, or exam is strictly prohibited. Be considerate and turn your cell phones off during class. If your cell phone makes any noise during class, then you will be asked to stand in front of the class and sing “I’m a little teapot” or similar song.

Academic Integrity: Students are encouraged to work together on homework and labs. I found working in small groups beneficial to increased comprehension of the material. THE NGCSU integrity code is: “On my honor, I will not lie, cheat, steal, plagiarize, evade the truth or tolerate those who do.” **ANY VIOLATION OF ANY OF THESE RULES, LIKE CHEATING OR PLAGARIZISM, WILL RESULT IN BEING FAILED IN THE COURSE, REPORTED TO THE V.P. FOR ACADEMIC AFFAIRS, THE DEAN OF STUDENTS AND THE ACADEMIC INTEGRITY COUNCIL, WHICH MAY ULTIMATELY RESULT IN SUSPENSION OR EXPULSION.**

Tenative Lecture Schedule (NOTICE: subject to change)

Date	Topic	Reading Pages
8/20	Introduction and Chapter 1	1-31
8/22	Chapter 1 – Chemistry: Matter and Measurement	1-31
8/25	Chapter 1 – Chemistry: Matter and Measurement	1-31
8/26	Drop/Add Ends	
8/27	Chapter 1 – Chemistry: Matter and Measurement	1-31
8/29	Chapter 2 – Atoms, Molecules, Ions	33-67
9/1	Labor Day Holiday- -No classes	
9/3	Chapter 2 – Atoms, Molecules, Ions	33-67
9/5	Chapter 2 – Atoms, Molecules, Ions	33-67
9/8	Chapter 2 – Atoms, Molecules, Ions	33-67
9/10	Chapter 3 – Formulas, Equations, and Moles	69-106
9/12	Exam 1 (Chapters 1 & 2)	
9/15	Chapter 3 – Formulas, Equations, and Moles	69-106
9/17	Chapter 3 – Formulas, Equations, and Moles	69-106
9/19	Chapter 4 – Reactions in Aqueous Solution	107-146
9/22	Chapter 4 – Reactions in Aqueous Solution	107-146
9/24	Chapter 4 – Reactions in Aqueous Solution	107-146
9/26	Chapter 4 – Reactions in Aqueous Solution	107-146
9/29	Chapter 4 – Reactions in Aqueous Solution	107-146
10/1	Chapter 5 – Periodicity and Atomic Structure	147-184
10/3	Exam 2 (Chapters 3 & 4)	
10/6	Chapter 5 – Periodicity and Atomic Structure	147-184
10/7	Last Day to Drop with a W	
10/8	Chapter 5 – Periodicity and Atomic Structure	147-184
10/10	Fall Break- -No classes	
10/13	Chapter 5 – Periodicity and Atomic Structure	147-184
10/15	Chapter 5 – Periodicity and Atomic Structure	147-184
10/17	Chapter 6 – Ionic Bonds and Some Main Group Chemistry	185-222
10/20	Chapter 6 – Ionic Bonds and Some Main Group Chemistry	185-222
10/22	Chapter 6 – Ionic Bonds and Some Main Group Chemistry	185-222
10/24	Chapter 7 – Covalent Bonds and Molecular Structure	223-272
10/27	Chapter 7 – Covalent Bonds and Molecular Structure	223-272
10/29	Chapter 7 – Covalent Bonds and Molecular Structure	223-272
10/31	Chapter 7 – Covalent Bonds and Molecular Structure	223-272
11/3	Chapter 7 – Covalent Bonds and Molecular Structure	223-272
11/5	Chapter 8 – Thermochemistry: Chemical Energy	273-314
11/7	Exam 3 (Chapters 5, 6 & 7)	
11/10	Chapter 8 – Thermochemistry: Chemical Energy	273-314
11/12	Chapter 8 – Thermochemistry: Chemical Energy	273-314
11/14	Chapter 9 – Gases: Their Properties and Behavior	315-352
11/17	Chapter 9 – Gases: Their Properties and Behavior	315-352
11/19	Chapter 9 – Gases: Their Properties and Behavior	315-352
11/21	Chapter 9 – Gases: Their Properties and Behavior	315-352
11/24	Exam 4 (Chapters 8 & 9)	
11/26-11/28	THANKSGIVING BREAK – No Classes	
12/1	Final Exam Question and Answer	
12/2	Last day of Classes	
12/3	Academic Review Day	
12/5	Final Exam – 10:30-12:30	